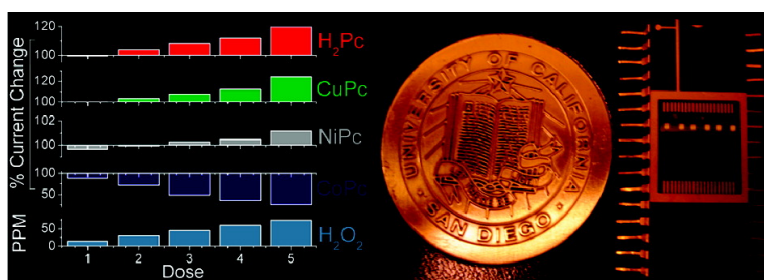


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Selective Detection of Vapor Phase Hydrogen Peroxide with Phthalocyanine Chemiresistors

Forest I. Bohrer, Corneliu N. Colesniuc, Jeongwon Park, Ivan K. Schuller, Andrew C. Kummel, and William C. Trogler*

Department of Chemistry and Biochemistry, University of California, San Diego, 9500 Gilman Drive, La Jolla, California 92093

Received November 14, 2007; E-mail: wtrogler@ucsd.edu

Vapor phase monitoring of peroxides is a topic of critical importance; recent incidents in England and Germany involving peroxide based explosives have made peroxide detection crucial to counterterrorism efforts.¹ Due to its widespread use and toxicity (OSHA PEL = 1 ppm), vapor phase monitoring of hydrogen peroxide (H₂O₂) is also an important industrial health issue.² Standard methods of H₂O₂ detection are limited to solid and liquid samples.³ Therefore, there is an urgent need for rapid vapor phase detection of H₂O₂ with a simple electronic sensor.

Phthalocyanines, both metalated (MPcs; M = p-, d-, and f-block elements) and metal-free (H₂Pc), are organic p-type semiconductors that are used as chemiresistive gas sensors.⁴ MPc conductivities are increased by oxidant gases that generate charge carriers (holes) and are decreased by electron-donating gases which trap charge carriers.^{5–6} MPcs act as catalysts for solution phase electrochemical oxidation and reduction of H₂O₂;⁷ however, MPcs have not been explored as vapor phase H₂O₂ sensors.

Chemiresistors using 50 nm thick films of MPcs (M = Co, Ni, Cu, and H₂) are examined as sensors for vapor phase H₂O₂ detection under practical conditions (i.e., in the presence of ambient humidity). Selectivity is gained from the ability of H₂O₂ to cause current losses in CoPc sensors and current gains in NiPc, CuPc, and H₂Pc sensors; this is the first example of contrasting analyte redox behavior dependent on the MPc metal center in a chemiresistor. The sensors were fabricated and tested as reported previously.⁶ Mass flow controllers were combined with impinger flasks to introduce known concentrations of H₂O₂ from a 27.1% ± 2.0% H₂O₂(aq) solution into the testing chamber.⁸

Sensor responses were recorded as time-dependent current plots at constant voltage and temperature (8 V, 50 ± 0.1 °C). There are two temporal components to sensor response and recovery, a fast adsorption/desorption process followed by a slow saturation/desaturation of analyte on the sensing surface.⁶ The fast (kinetic) component of the response has been attributed to binding of analyte at oxygen-free metal sites, and the slow (saturation) component, to competitive displacement of bound oxygen. At room temperature, the crossover point between the fast and slow components occurs near 5 min. To probe both temporal components, sensors were dosed with identical concentrations of 27% H₂O₂(aq) in both the kinetic regime (5 min doses with 90 min recoveries) and the saturation regime (30 min doses with 90 min recoveries). The doses consisted of 45 ppm of H₂O₂ accompanied by 4950 ppm of water (relative humidity RH = 17%). In parallel experiments, the sensors were dosed at constant humidity (RH = 17%) in order to distinguish the sensor response of H₂O₂ from that of water. These data are given in Figure 1.

CoPc responds with current losses to each dose of H₂O₂, with or without constant humidity, in both the kinetic and saturation regimes. Conversely, NiPc, CuPc, and H₂Pc respond to H₂O₂ with current gains in the presence or absence of constant humidity. In

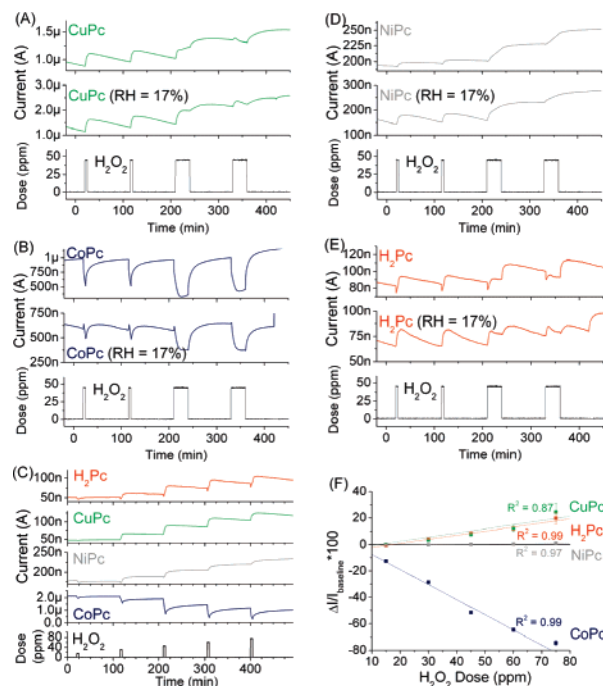


Figure 1. (A) CuPc (green traces) sensing data (50 °C, 8 V) on exposure to 27% H₂O₂(aq) doses (black trace) in the presence and absence of constant humidity. RH = relative humidity. (B) CoPc (blue traces) sensing data. (C) Sensing data for MPcs (M = Co, Ni, Cu, H₂) exposed to varied concentration doses of 27% H₂O₂(aq). (D) NiPc (gray traces) sensing data. (E) H₂Pc (red traces) sensing data. (F) Quantitative sensor response data plotted for each dose in (C).

the saturation regime, all sensors reach a maximum response within 10 min. MPc sensors generally exhibit current losses on exposure to water vapor.^{5–6} The NiPc, CuPc, and H₂Pc responses to 27% H₂O₂(aq) are therefore consistent with competing redox effects of H₂O₂ oxidizing the film and water reducing the film. The H₂O₂–MPc film interaction is dominant even in the presence of a concentration of water 100 times greater than H₂O₂. Strong oxidants cause current gains in all CoPc, NiPc, CuPc, and H₂Pc films, while electron donors cause current losses in these same films; H₂O₂ is unique in showing current losses for CoPc and current gains in other MPc films. This is illustrated in Figure 2, which displays sensor responses of CoPc, CuPc, and H₂Pc to selected electron donors⁶ and to H₂O₂. Thus a CoPc sensor could be paired with a CuPc/H₂Pc sensor for a peroxide-specific array.

Sensor responses of MPcs exhibit first-order kinetics and therefore are linear with analyte concentration for physisorption and chemisorption interactions.⁹ H₂O₂ adsorption kinetics were analyzed by varying the dose concentration (15, 30, 45, 60, and 75 ppm H₂O₂, Figure 1C). Quantitative sensor responses (Figure 1F,

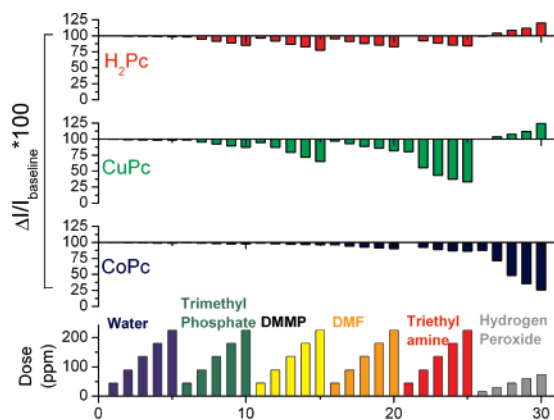


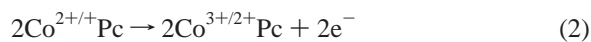
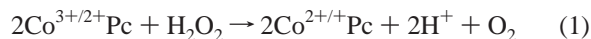
Figure 2. Sensor responses ($\Delta I/I_{\text{baseline}} \times 100$) for CoPc (blue), CuPc (green), and H₂Pc (red) exposed to various electron donors and H₂O₂.

Table 1. MPc Sensitivities to H₂O₂ and Estimated Detection Limits

MPc	H ₂ O ₂ sensitivity (% ppm ⁻¹) × 10 ⁻²	H ₂ O ₂ detection limit (ppm)
CoPc	-103.2 ± 3.3	0.05
NiPc	2.3 ± 0.9	40.1
CuPc	31.2 ± 8.3	12.2
H ₂ Pc	27.6 ± 5.2	11.7

$\Delta I/I_{\text{baseline}} \times 100$) for CoPc, NiPc, CuPc, and H₂Pc in the kinetic regime are linear with concentration. MPc sensitivities to H₂O₂ are defined as the slope (% ppm⁻¹) of the sensor response versus dose concentration (Figure 1F, Table 1). Conversely, in the saturation regime, the sensor responses are not linear with H₂O₂ concentration; instead, the responses depend on dosing history (Figure S1). This unusual behavior is consistent with a chemical or electrochemical reaction modifying the sensor film, rather than simple physisorption or chemisorption.

CoPc is a potent catalyst for H₂O₂ oxidation and reduction, and the mechanism of H₂O₂ decomposition in solution has been previously explored.¹⁰ It is inferred from the large current losses observed for CoPc in Figure 1C that cobalt-catalyzed oxidation of H₂O₂ on the surface of the CoPc film is occurring (with concurrent film reduction), leading to hole trapping and loss of current. The reaction can be written as follows:¹⁰



In contrast, voltammetric studies of MPcs with nonredox-active metal centers (M = Ni, Cu, and H₂) revealed no electrocatalytic behavior in the presence of H₂O₂,¹¹ but EPR studies of optically excited ZnPc solutions revealed the presence of •OH radicals, suggesting homolytic cleavage of H₂O₂.¹² The presence of small amounts of hydroxyl radicals generated during dosing may account for the partial oxidation of NiPc, CuPc, and H₂Pc films, leading to the observed sensor responses as radical reactions (eq 4) or as charge-transfer processes (eq 5):¹²



MPc chemiresistive sensors offer advantages over existing commercial vapor phase H₂O₂ detection methods, which suffer from

drawbacks such as cross-reactive interferents and complex instrumentation.¹³ Detection limits of the cited commercial H₂O₂ sensors range from 0.1 to 1 ppm. Detection limits for the chemiresistive MPc sensors examined in this study were calculated from the MPc sensitivities (Table 1) at a signal-to-noise ratio of 3. Due to its high catalytic activity, CoPc is the most sensitive material for detecting H₂O₂, with a detection limit of 50 ppb. These limits may be further improved by incorporation of a preconcentrator or an ultrathin ChemFET geometry.¹⁴ Detection of peroxide-based explosives (TATP, HMTD) may be achieved by liberation of H₂O₂ from these explosives by UV light,³ followed by detection with the MPc sensor array.

In summary we have shown that H₂O₂ vapors may be selectively detected by contrasting oxidation/reduction behavior in nanoscale chemiresistive films of MPcs. This study presents the first example of MPc vapor sensors showing contrasting current responses to the same analyte based on the specific metal center, which may be attributed to electrocatalytic processes occurring in the sensor film. Therefore, differential response analysis can be used to selectively identify the presence of H₂O₂ through the combination of only two sensors (e.g., CoPc and CuPc) with opposite sensor responses (Figure 2). This study represents a new approach to selective analyte detection: catalytic redox contrast in a sensor array.

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Supporting Information Available: Fabrication and experimental techniques and saturation regime data (Figure S1). This material is available free of charge via the Internet at <http://pubs.acs.org>.

References

- (a) Bellamy, A. J. *J. Forensic Sci.* **1999**, *44*, 603–608. (b) *Report of the Official Account of the Bombings in London on 7th July 2005*. HC series, 2005–06 1087; House of Commons, Parliament, Great Britain; Stationary Office, London, 2006. (c) Landler, Mark. German police arrest 3 in terrorist plot. *New York Times* 09/06/2007; http://www.nytimes.com/2007/09/06/world/europe/06_germany.html.
- Occupational Safety and Health Administration Website, Chemical Sampling Information: Hydrogen Peroxide. http://www.osha.gov/dts/chemicalsampling/data/CH_246600.html (accessed May 2007).
- (a) Schulte-Ladbeck, R.; Edelmann, A.; Quintas, G.; Lendl, B.; Karst, U. *Anal. Chem.* **2006**, *78*, 8150–8155. (b) Schulte-Ladbeck, R.; Kolla, P.; Karst, U. *Analyst* **2002**, *127*, 1152–1154.
- Snow, A. W.; Barger, W. R. *Phthalocyanine Films in Chemical Sensors. Phthalocyanines: Properties and Applications*; Lever, A. B. P., Ed.; Wiley and Sons: New York, 1989; Vol. 1, p 341.
- (a) Liu, C. J.; Hsieh, J. C.; Ju, Y. H. *J. Vac. Sci. Technol., A* **1996**, *14* (3), 753–756. (b) Bora, M.; Schut, D.; Baldo, M. A. *Anal. Chem.* **2007**, *79*, 3298–3303.
- Bohrer, F. I.; Sharoni, A.; Colesniuc, C.; Park, J.; Schuller, I. K.; Kummel, A. C.; Trogler, W. C. *J. Am. Chem. Soc.* **2007**, *129*, 5640–5646.
- (a) Wang, J.; Golden, T.; Li, R. *Anal. Chem.* **1988**, *60*, 1642–1645. (b) Qi, X.; Baldwin, R. P. *Electroanalysis* **1993**, *5*, 547–554.
- Manatt, S. L.; Manatt, M. R. *Chem.—Eur. J.* **2004**, *10*, 6540–6557.
- Tongpool, R.; Yoriya, S. *Thin Solid Films* **2005**, *477*, 148–152.
- (a) Gilmartin, M. A. T.; Ewen, R. J.; Hart, J. P.; Honeybourne, C. L. *Electroanalysis* **1995**, *7*, 547–555. (b) Tarasevich, M. R.; Zakharkin, G. I. *React. Kinet. Catal. Lett.* **1977**, *6*, 77–82.
- Gilmartin, M. A. T.; Ewen, R. J.; Hart, J. P. *J. Electroanal. Chem.* **1996**, *401*, 127–137.
- Gantchev, T. G.; Sharman, W. M.; Lier, J. E. *Photochem. Photobiol.* **2003**, *77*, 469–479.
- (a) Dräger Inc. Safety Website, Dräger Safety Short Term Detector Tubes: Hydrogen Peroxide. http://www.draeger.com/ST/internet/pdf/US/detection/brochure_tubes_e.pdf (accessed July 2007). (b) McVey, I. F. Non-dispersive mid-infrared sensor for vaporized hydrogen peroxide. U.S. Patent 7,157,045, July 9, 2002. (c) Dräger Inc. Safety Website, DrägerSensor H₂O₂ LC Chemical Sensor Data Sheet. http://www.draeger.com/ST/internet/pdf/Master/En/gt/9023492_h2o2lc_d_e.pdf (accessed July 2007).
- (a) Grate, J. W.; Anheier, N. C.; Baldwin, D. L. *Anal. Chem.* **2005**, *77*, 1867–1875. (b) Yang, R. D.; Gredig, T.; Colesniuc, C. N.; Park, J.; Schuller, I. K.; Trogler, W. C.; Kummel, A. C. *Appl. Phys. Lett.* **2007**, *90*, 263506/1–3.

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